

Quiz for Video 15 – Intermolecular Forces

- Only polar molecules will experience dipole-dipole forces.
 - True
 - False
- Which type of intermolecular force is felt by ALL matter.
 - London dispersion
 - Dipole-dipole
 - Hydrogen-bonding
 - Ionic bonding
- Which of the following physical properties is NOT affected by intermolecular forces.
 - Color
 - Viscosity
 - Surface Tension
 - Boiling Point
- If molecule A has a stronger net dipole than molecule B, what should we expect about the boiling point of the two molecules? (Assume molar mass is approximately the same).
 - A will have a higher boiling point
 - B will have a higher boiling point
 - A and B will have the same boiling point
 - Both will always be solid
- The presence of which of the following covalent bonds would result in a molecule exhibiting hydrogen bonding?
 - H-Cl
 - H-O
 - H-C
 - H-S
- As we increase the external pressure around a compound, what state is generally favored? (note: water is a major exception!)
 - gas
 - liquid
 - solid
- Molecules with stronger intermolecular forces will have boiling points that are:
 - At higher Temperatures
 - At lower Temperatures
 - At room Temperature
 - Not enough information
- Identify the strongest intermolecular force present in ammonia (NH_3).
 - London dispersion
 - Dipole-dipole
 - Hydrogen-bonding
 - Ionic bonding
- Identify the strongest intermolecular force present in ethane (C_2H_6).
 - London dispersion
 - Dipole-dipole
 - Hydrogen-bonding
 - Ionic bonding
- Identify the strongest intermolecular force present in difluoromethane (CH_2F_2).
 - London dispersion
 - Dipole-dipole
 - Hydrogen-bonding
 - Ionic bonding