

# Introduction for Module 6 – Titration

Textbook: [Open Stax Chemistry 2e](#)

Suggested Reading: Chapter 14.7

Learning Objectives:

- Interpret titration curves for strong and weak acid-base systems
- Compute sample pH at important stages of a titration
- Explain the function of acid-base indicators

Captions and Attributions:

- 1) As an acidic sample (analyte) has strong base of known concentration (titrate) added the pH will rise in a predictable way, which can be used to quantify the analyte originally present. [Figure 14.18 \(a\) The titration curve](#) by [OpenStax](#) is licensed under [CC BY 4.0](#).
- 2) Titration is a laboratory method for determining the concentration of an analyte using a titrant of known concentration. [Titration - laboratory method](#) by [Wikimedia Commons](#) is licensed under [CC BY SA 4.0](#)



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