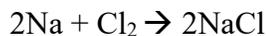


Quiz for Video 5 – Limiting Reactants

1. A chemical equation is like a recipe.
 - a. True
 - b. False
2. What determines which reactant is limiting?
 - a. The one with smaller mass
 - b. The one that makes less product
 - c. The one with smaller stoichiometric coefficients
 - d. The one with smaller molar mass
3. With 8 slices of bread and 5 slices of cheese, how many sandwiches would I make, following the example in the video?
 - a. 4
 - b. 5
 - c. 6
 - d. 8

4. With 8 slices of bread and 5 slices of cheese, which is the “limiting reactant,” following the example in the video?
 - a. Bread
 - b. Cheese
 - c. Both
 - d. Neither

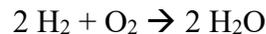
5. With 8 slices of bread and 5 slices of cheese, what “reactant” will remain, following the example in the video?
 - a. 2 slices of bread
 - b. 1 slice of cheese
 - c. Nothing remains
 - d. 5 sandwiches
6. Which is the limiting reactant in the reaction below?



- a. Na
- b. Cl
- c. Impossible to tell

7. Do we need to balance equations before considering limiting reactant?
 - a. Yes
 - b. No

8. For the reaction forming water below:



How much water will be formed from 4 moles H_2 and 1 mole O_2 ?

- a. 1 mole H_2O
- b. 2 moles H_2O
- c. 3 moles H_2O
- d. 4 moles H_2O

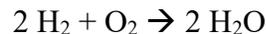
9. For the reaction forming water below:



How much water will be formed from 10.0g H_2 and 20.0g O_2 ?

- a. 11.3g H_2O
- b. 22.5g H_2O
- c. 45.0g H_2O
- d. 89.3g H_2O

10. For the reaction forming water below:



How much of the excess reactant will remain after combining 10.0g H_2 and 20.0g O_2 ?

- a. 2.5g O_2
- b. 7.5g H_2
- c. 10g O_2
- d. 15g O_2