

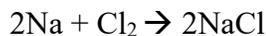
Quiz for Video 5 – Limiting Reactants

1. A chemical equation is like a recipe.
a. True
b. False
2. What determines which reactant is limiting?
a. The one with smaller mass
b. The one that makes less product
c. The one with smaller stoichiometric coefficients
d. The one with smaller molar mass
3. With 8 slices of bread and 5 slices of cheese, how many sandwiches would I make, following the example in the video?
a. 4
b. 5
c. 6
d. 8

4. With 8 slices of bread and 5 slices of cheese, which is the “limiting reactant,” following the example in the video?
a. Bread
b. Cheese
c. Both
d. Neither

5. With 8 slices of bread and 5 slices of cheese, what “reactant” will remain, following the example in the video?
a. 2 slices of bread
b. 1 slice of cheese
c. Nothing remains
d. 5 sandwiches

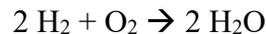
6. Which is the limiting reactant in the reaction below?



- a. Na
b. Cl
c. Impossible to tell

7. Do we need to balance equations before considering limiting reactant?
a. Yes
b. No

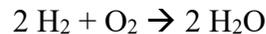
8. For the reaction forming water below:



How much water will be formed from 4 moles H_2 and 1 mole O_2 ?

- a. 1 mole H_2O
b. 2 moles H_2O
c. 3 moles H_2O
d. 4 moles H_2O

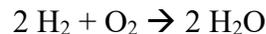
9. For the reaction forming water below:



How much water will be formed from 10.0g H_2 and 20.0g O_2 ?

- a. 11.3g H_2O
b. 22.5g H_2O
c. 45.0g H_2O
d. 89.3g H_2O

10. For the reaction forming water below:



How much of the excess reactant will remain after combining 10.0g H_2 and 20.0g O_2 ?

- a. 2.5g O_2
b. 7.5g H_2
c. 10g O_2
d. 15g O_2