

## Quiz for Video 5 – Limiting Reactants

1. A chemical equation is like a recipe.

- a. True
- b. False

2. What determines which reactant is limiting?

- a. The one with smaller mass
- b. The one that makes less product
- c. The one with smaller stoichiometric coefficients
- d. The one with smaller molar mass

3. With 8 slices of bread and 5 slices of cheese, how many sandwiches would I make, following the example in the video?

- a. 4
- b. 5
- c. 6
- d. 8

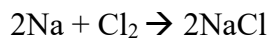
4. With 8 slices of bread and 5 slices of cheese, which is the “limiting reactant,” following the example in the video?

- a. Bread
- b. Cheese
- c. Both
- d. Neither

5. With 8 slices of bread and 5 slices of cheese, what “reactant” will remain, following the example in the video?

- a. 2 slices of bread
- b. 1 slice of cheese
- c. Nothing remains
- d. 5 sandwiches

6. Which is the limiting reactant in the reaction below?

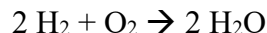


- a. Na
- b. Cl
- c. Impossible to tell

7. Do we need to balance equations before considering limiting reactant?

- a. Yes
- b. No

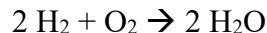
8. For the reaction forming water below:



How much water will be formed from 4 moles  $\text{H}_2$  and 1 mole  $\text{O}_2$ ?

- a. 1 mole  $\text{H}_2\text{O}$
- b. 2 moles  $\text{H}_2\text{O}$
- c. 3 moles  $\text{H}_2\text{O}$
- d. 4 moles  $\text{H}_2\text{O}$

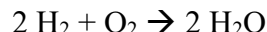
9. For the reaction forming water below:



How much water will be formed from 10.0g  $\text{H}_2$  and 20.0g  $\text{O}_2$ ?

- a. 11.3g  $\text{H}_2\text{O}$
- b. 22.5g  $\text{H}_2\text{O}$
- c. 45.0g  $\text{H}_2\text{O}$
- d. 89.3g  $\text{H}_2\text{O}$

10. For the reaction forming water below:



How much of the excess reactant will remain after combining 10.0g  $\text{H}_2$  and 20.0g  $\text{O}_2$ ?

- a. 2.5g  $\text{O}_2$
- b. 7.5g  $\text{H}_2$
- c. 10g  $\text{O}_2$
- d. 15g  $\text{O}_2$