

Quiz for Video 6 – Redox Reactions

1. What does the term “redox” stand for in redox reaction?
 - a. Re-dosing
 - b. Reduction-Oxidation
 - c. Red-Oxen
 - d. Recombining-Oxygen
2. A redox reaction represents the flow of:
 - a. Electrons (e-)
 - b. Protons (H+)
 - c. Positrons (e+)
 - d. Atoms
3. Reduction is associated with the ____ of electrons.
 - a. Losing
 - b. Gaining
 - c. Sharing
 - d. Preservation
4. Oxidation is associated with the ____ of electrons.
 - a. Losing
 - b. Gaining
 - c. Sharing
 - d. Preservation
5. Why do we refer to “oxidation number” instead of charge?
 - a. We don't need to
 - b. The real charge is distributed across multiple atoms
 - c. Because we are losing or gaining electrons
6. What is the oxidation number of Oxygen (generally) when in a compound?
 - a. 0
 - b. +2
 - c. -2
 - d. -1
7. Why do elements always have an oxidation number of 0?
 - a. Because they are neutral
 - b. Because they are diatomic
 - c. Because they are in the gas state
 - d. They don't necessarily
8. In a certain reaction, an atom goes from an oxidation number of +2 to an oxidation number of -4. This atom is:
 - a. Reduced
 - b. Oxidized
 - c. Eliminated
 - d. Unchanged
9. True or false: oxidation and reduction always occur together.
 - a. True
 - b. False
10. In the reaction below, which species is the oxidizing agent?
$$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 3\text{H}_2\text{O}$$
 - a. C_3H_8 , because it is reduced
 - b. C_3H_8 , because it is oxidized
 - c. O_2 , because it is reduced
 - d. O_2 , because it is oxidized