

Introduction for Module 10 – Solubility

Textbook: [Open Stax Chemistry 2e](#)

Suggested Reading: Chapter 15.1

Learning Objectives:

- **Write chemical equations and equilibrium expressions representing solubility equilibria**
- **Carry out equilibrium computations involving solubility, equilibrium expressions, and solute concentrations**

Captions and Attributions:

- 1) Silver chloride is considered “insoluble” but still does dissociate a very small amount. The equilibrium of this reaction is determined by the concentration of dissociated ions. [Figure 15.2 Silver chloride is a sparingly soluble ionic solid](#) by [Open Stax](#) is [licensed under CCBY 4.0](#).



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