1. Predict the products for the following reactions. Don’t forget states of matter!

MgCl2(aq) + NaOH(aq) 🡪

H2SO4(aq) + Pb(OH)4(s) 🡪

1. What type of reaction would you classify the reactions in part A? More than one may apply!
2. Combustion of hydrocarbons is a common redox reaction. State the oxidation numbers for all atoms in each of the reaction below, where the first shows the combustion of ethane (a component of natural gas) and the other shows the combustion of ethanol (a common gasoline additive).

C2H6(g) + 3.5 O2(g) 🡪 3 H2O(g) + 2 CO2(g)

C2H6O(g) + 3 O2(g) 🡪 3 H2O(g) + 2 CO2(g)

1. State the role of oxygen in each reaction above. Does this help to describe why we call losing electrons “oxidation”?