

## Quiz for Video 6 – Redox Reactions

1. What does the term “redox” stand for in redox reaction?
  - a. Re-dosing
  - b. Reduction-Oxidation
  - c. Red-Oxen
  - d. Recombining-Oxygen
2. A redox reaction represents the flow of:
  - a. Electrons (e-)
  - b. Protons (H+)
  - c. Positrons (e+)
  - d. Atoms
3. Reduction is associated with the \_\_\_\_ of electrons.
  - a. Losing
  - b. Gaining
  - c. Sharing
  - d. Preservation
4. Oxidation is associated with the \_\_\_\_ of electrons.
  - a. Losing
  - b. Gaining
  - c. Sharing
  - d. Preservation
5. Why do we refer to “oxidation number” instead of charge?
  - a. We don't need to
  - b. The real charge is distributed across multiple atoms
  - c. Because we are losing or gaining electrons
6. What is the oxidation number of Oxygen (generally) when in a compound?
  - a. 0
  - b. +2
  - c. -2
  - d. -1
7. Why do elements always have an oxidation number of 0?
  - a. Because they are neutral
  - b. Because they are diatomic
  - c. Because they are in the gas state
  - d. They don't necessarily
8. In a certain reaction, an atom goes from an oxidation number of +2 to an oxidation number of -4. This atom is:
  - a. Reduced
  - b. Oxidized
  - c. Eliminated
  - d. Unchanged
9. True or false: oxidation and reduction always occur together.
  - a. True
  - b. False
10. In the reaction below, which species is the oxidizing agent?
$$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 3\text{H}_2\text{O}$$
  - a.  $\text{C}_3\text{H}_8$ , because it is reduced
  - b.  $\text{C}_3\text{H}_8$ , because it is oxidized
  - c.  $\text{O}_2$ , because it is reduced
  - d.  $\text{O}_2$ , because it is oxidized